

Shared Instructor and Student Functionality

Last Modified on 08/30/2023 3:33 pm EDT

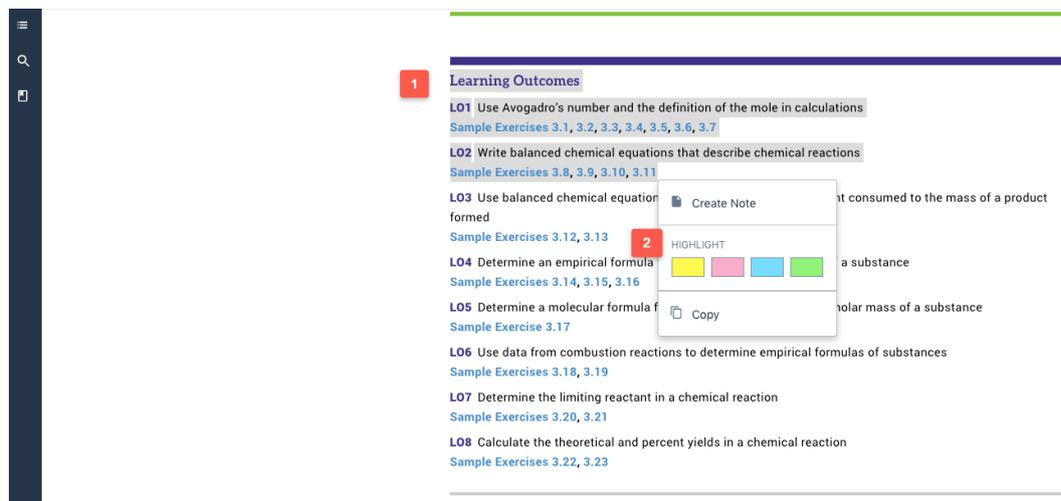
The Norton Ebook Reader has functionality that is common to both instructors and students. This page details the highlighting, annotation, bookmarking, audio narration, and printing and copying capabilities that are available to both instructors and students.

Hide All Answers

How do I highlight text?

Personal Highlights

To highlight text, use the cursor to select the text you would like to highlight and the **Context Menu** will appear. Select the color in which you would like the text highlighted: yellow, pink, blue, or green.



The screenshot displays a vertical sidebar on the left with icons for a menu, search, and a document. The main content area shows a list of Learning Outcomes (LO1-LO8) under the heading "Learning Outcomes". A red box with the number "1" is positioned to the left of the LO1 text. A red box with the number "2" is positioned to the left of the "HIGHLIGHT" option in the context menu. The context menu is open over the text "Determine an empirical formula" in LO4, showing options for "Create Note", "HIGHLIGHT" (with color selection buttons for yellow, pink, blue, and green), and "Copy".

Removing Highlights

To remove a highlight, use the cursor to select the text from which you would like to remove the highlighting and the **Context Menu** will appear.

Click **Delete Highlight**

Chemistry student123@mailinator.com Page 274

6: Properties of Gases: The Air We Br... > 6.1 Air: An Invisible Necessity

anesthesiologists in a hospital operating room constantly monitor levels of oxygen and carbon dioxide in the blood. The management of the delicate balance of gases entering and leaving a patient can mean the difference between a normal recovery and an irreversible coma.

We have seen how dissolved compounds react in aqueous solution. Chemical reactions also take place in the gas phase, and gases are intimately involved in chemical reactions in living systems as well as in the material world. Most life in our biosphere requires oxygen. Insects, birds, mammals, plants, and even underwater organisms need O_2 to metabolize nutrients.

1 How do gases differ from solids and liquids? Gases have neither definite volumes nor definite shapes; they expand to occupy the entire volume of their container and assume the container's shape. Under everyday conditions, other properties also distinguish gases from liquids and solids:

1. Unlike the volume occupied by a liquid or solid, the volume occupied by a gas changes significantly with pressure. If we carry an inflated balloon from sea level (0 m) to the top of a 1600-m mountain, the balloon volume increases by about 20%. The volume of a liquid or solid is unchanged under these conditions.
2. The volume of a gas changes with temperature. For example, the volume of a balloon filled with room-temperature air decreases when the balloon is taken outside on a cold winter's day. A temperature decrease from 20°C to 0°C leads to a volume decrease of about 7%, whereas the volume of a liquid or solid remains practically unchanged by this modest temperature change.
3. Gases are **miscible**, which means they can be mixed in any proportion (unless they chemically

Type your annotation into the text field and click the **Save** button save your annotation.

Chemistry student123@mailinator.com Page 274

6: Properties of Gases: The Air We Br... > 6.1 Air: An Invisible Necessity

anesthesiologists in a hospital operating room constantly monitor levels of oxygen and carbon dioxide in the blood. The management of the delicate balance of gases entering and leaving a patient can mean the difference between a normal recovery and an irreversible coma.

We have seen how dissolved compounds react in aqueous solution. Chemical reactions also take place in the gas phase, and gases are intimately involved in chemical reactions in living systems as well as in the material world. Most life in our biosphere requires oxygen. Insects, birds, mammals, plants, and even underwater organisms need O_2 to metabolize nutrients.

How do gases differ from solids and liquids? Gases have neither definite volumes nor definite shapes; they expand to occupy the entire volume of their container and assume the container's shape. Under everyday conditions, other properties also distinguish gases from liquids and solids:

1. Unlike the volume occupied by a liquid or solid, the volume occupied by a gas changes significantly with pressure. If we carry an inflated balloon from sea level (0 m) to the top of a 1600-m mountain, the balloon volume increases by about 20%. The volume of a liquid or solid is unchanged under these conditions.
2. The volume of a gas changes with temperature. For example, the volume of a balloon filled with room-temperature air decreases when the balloon is taken outside on a cold winter's day. A temperature decrease from 20°C to 0°C leads to a volume decrease of about 7%, whereas the volume of a liquid or solid remains practically unchanged by this modest temperature change.
3. Gases are **miscible**, which means they can be mixed in any proportion (unless they chemically react with one another). A hospital patient experiencing respiratory difficulties may be given a mixture of nitrogen and oxygen in which the proportion of oxygen is much higher than its proportion in air. Alternatively, a scuba diver may leave the ocean surface with a tank of air

Click on the **Notebook page icon** to view notes in the Notebook

Chemistry student123@mailinator.com Page 274

6: Properties of Gases: The Air We Br... > 6.1 Air: An Invisible Necessity

anesthesiologists in a hospital operating room constantly monitor levels of oxygen and carbon dioxide in the blood. The management of the delicate balance of gases entering and leaving a patient can mean the difference between a normal recovery and an irreversible coma.

We have seen how dissolved compounds react in aqueous solution. Chemical reactions also take place in the gas phase, and gases are intimately involved in chemical reactions in living systems as well as in the material world. Most life in our biosphere requires oxygen. Insects, birds, mammals, plants, and even underwater organisms need O_2 to metabolize nutrients.

How do gases differ from solids and liquids? Gases have neither definite volumes nor definite shapes; they expand to occupy the entire volume of their container and assume the container's shape. Under everyday conditions, other properties also distinguish gases from liquids and solids:

1. Unlike the volume occupied by a liquid or solid, the volume occupied by a gas changes significantly with pressure. If we carry an inflated balloon from sea level (0 m) to the top of a 1600-m mountain, the balloon volume increases by about 20%. The volume of a liquid or solid is unchanged under these conditions.
2. The volume of a gas changes with temperature. For example, the volume of a balloon filled with room-temperature air decreases when the balloon is taken outside on a cold winter's day. A temperature decrease from 20°C to 0°C leads to a volume decrease of about 7%, whereas the volume of a liquid or solid remains practically unchanged by this modest temperature change.
3. Gases are **miscible**, which means they can be mixed in any proportion (unless they chemically react with one another). A hospital patient experiencing respiratory difficulties may be given a mixture of nitrogen and oxygen in which the proportion of oxygen is much higher than its proportion in air. Alternatively, a scuba diver may leave the ocean surface with a tank of air containing a homogeneous mixture of 17% oxygen, 34% nitrogen, and 49% helium. In contrast, many liquids are immiscible, such as oil and water.
4. Gases are typically much less dense than liquids or solids. One indicator of this large difference is that gas densities are expressed in grams per liter but liquid densities are expressed in grams per milliliter. The density of dry air at 20°C at typical atmospheric pressure is 1.20 g/L, for example, whereas the density of liquid water under the same conditions is 1.00 g/mL—more than 800 times greater than the density of dry air.

These four observations about gases are consistent with the idea that the particles of a gas (be they molecules or atoms) are further apart than the particles in solids and liquids. The larger

How to Edit Annotations

Click the **notebook page icon**. The **Context Menu** will appear. Select **Edit Note**

Chemistry student123@mailinator.com

6: Properties of Gases: The Air We Br... > 6.1 Air: An Invisible Necessity Page 274

anesthesiologists in a hospital operating room constantly monitor levels of oxygen and carbon dioxide in the blood. The management of the delicate balance of gases entering and leaving a patient can mean the difference between a normal recovery and an irreversible coma.

We have seen how dissolved compounds react in aqueous solution. Chemical reactions also take place in the gas phase, and gases are intimately involved in chemical reactions in living systems as well as in the material world. Most life in our biosphere requires oxygen. Insects, birds, mammals, plants, and even underwater organisms need O₂ to metabolize nutrients.

How do gases differ from solids and liquids? Gases have neither definite volumes nor definite shapes; they expand to occupy the entire volume of their container and assume the container's shape. Under everyday conditions, other properties also distinguish gases from liquids and solids.

1. Unlike a liquid or solid, the volume occupied by a gas changes significantly with pressure. If we carry an inflated balloon from sea level (0 m) to the top of a 1600-m mountain, the balloon volume increases by about 20%. The volume of a liquid or solid is unchanged under these conditions.

2. The volume of a gas changes with temperature. For example, the volume of a balloon filled with room-temperature air decreases when the balloon is taken outside on a cold winter's day. A temperature decrease from 20°C to 0°C leads to a volume decrease of about 7%, whereas the volume of a liquid or solid remains practically unchanged by this modest temperature change.

3. Gases are **miscible**, which means they can be mixed in any proportion (unless they chemically react with one another). A hospital patient experiencing respiratory difficulties may be given a mixture of gases in which the proportion of oxygen is much higher than its proportion in air. Alternatively, a scuba diver may leave the ocean surface with a tank of air containing a homogeneous mixture of 17% oxygen, 34% nitrogen, and 49% helium. In contrast, many liquids are immiscible, such as oil and water.

4. Gases are typically much less dense than liquids or solids. One indicator of this large difference

After editing the note, select **Save**.

Chemistry student123@mailinator.com

6: Properties of Gases: The Air We Br... > 6.1 Air: An Invisible Necessity Page 274

anesthesiologists in a hospital operating room constantly monitor levels of oxygen and carbon dioxide in the blood. The management of the delicate balance of gases entering and leaving a patient can mean the difference between a normal recovery and an irreversible coma.

We have seen how dissolved compounds react in aqueous solution. Chemical reactions also take place in the gas phase, and gases are intimately involved in chemical reactions in living systems as well as in the material world. Most life in our biosphere requires oxygen. Insects, birds, mammals, plants, and even underwater organisms need O₂ to metabolize nutrients.

How do gases differ from solids and liquids? Gases have neither definite volumes nor definite shapes; they expand to occupy the entire volume of their container and assume the container's shape. Under everyday conditions, other properties also distinguish gases from liquids and solids.

1. Unlike the volume occupied by a liquid or solid, the volume occupied by a gas changes significantly with pressure. If we carry an inflated balloon from sea level (0 m) to the top of a 1600-m mountain, the balloon volume increases by about 20%. The volume of a liquid or solid is unchanged under these conditions.

2. The volume of a gas changes with temperature. For example, the volume of a balloon filled with room-temperature air decreases when the balloon is taken outside on a cold winter's day. A temperature decrease from 20°C to 0°C leads to a volume decrease of about 7%, whereas the volume of a liquid or solid remains practically unchanged by this modest temperature change.

3. Gases are **miscible**, which means they can be mixed in any proportion (unless they chemically react with one another). A hospital patient experiencing respiratory difficulties may be given a

How to Delete Annotations

1. Select the **notebook page icon** on the annotation that you want to delete
2. Click **Delete Highlight & Note**

Chemistry student123@mailinator.com

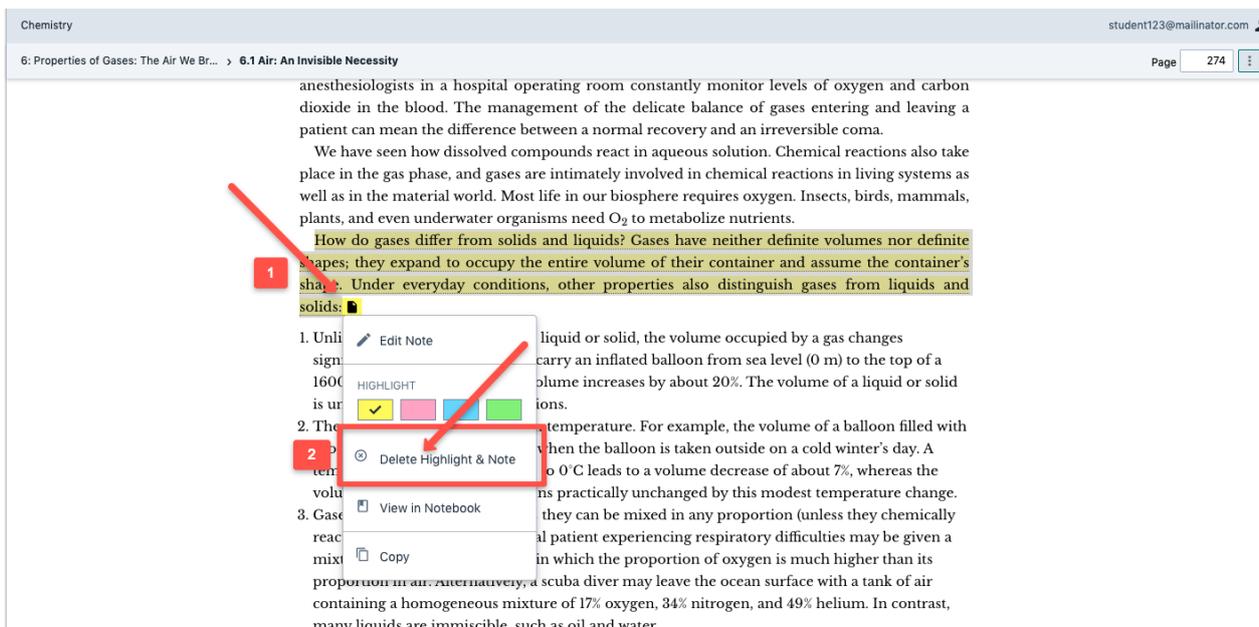
6: Properties of Gases: The Air We Br... > 6.1 Air: An Invisible Necessity Page 274

anesthesiologists in a hospital operating room constantly monitor levels of oxygen and carbon dioxide in the blood. The management of the delicate balance of gases entering and leaving a patient can mean the difference between a normal recovery and an irreversible coma.

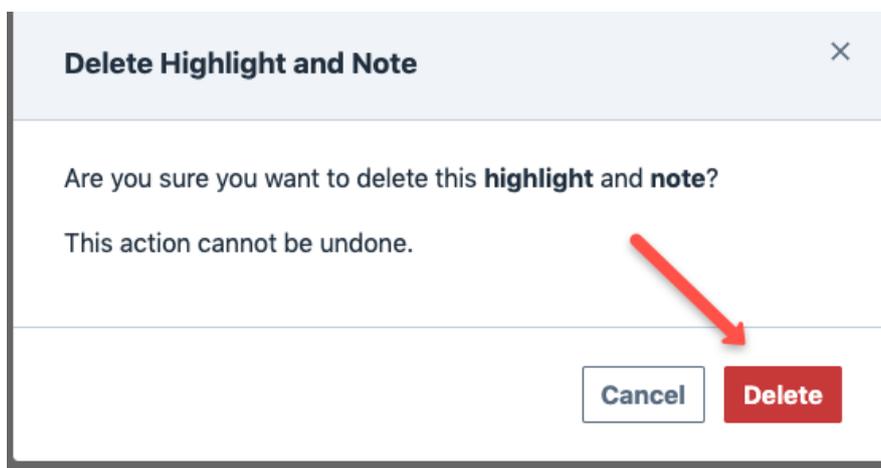
We have seen how dissolved compounds react in aqueous solution. Chemical reactions also take place in the gas phase, and gases are intimately involved in chemical reactions in living systems as well as in the material world. Most life in our biosphere requires oxygen. Insects, birds, mammals, plants, and even underwater organisms need O_2 to metabolize nutrients.

How do gases differ from solids and liquids? Gases have neither definite volumes nor definite shapes; they expand to occupy the entire volume of their container and assume the container's shape. Under everyday conditions, other properties also distinguish gases from liquids and solids:

- 1. Unlike liquids and solids, gases are compressible.
- 2. The volume of a gas changes with temperature. For example, the volume of a balloon filled with air at sea level (0 m) and 20°C increases by about 20% when the balloon is taken outside on a cold winter's day. A balloon filled with air at 0°C leads to a volume decrease of about 7%, whereas the volume of a liquid or solid is practically unchanged by this modest temperature change.
- 3. Gases can be mixed in any proportion (unless they chemically react). For example, a patient experiencing respiratory difficulties may be given a mixture of gases in which the proportion of oxygen is much higher than its proportion in air. Alternatively, a scuba diver may leave the ocean surface with a tank of air containing a homogeneous mixture of 17% oxygen, 34% nitrogen, and 49% helium. In contrast, many liquids are immiscible, such as oil and water.



Click the **Delete** button to confirm



To view a complete list of the highlights and annotations in your ebook, select the **Notebook icon** on the left of the page

Chemistry

1: Particles of Matter: Measurement and the Tools of Science

Page 2

↑ Previous:

1

Particles of Matter

Measurement and the Tools of Science

Chemistry

1: Particles of Matter: Measurement and the Tools of Science

↑ Previous:

1

Particles of Matter

Measurement and the Tools

1. This is the **total number** of notes and highlights
2. To **Edit** or **Delete** content select the 3 dots icon above the annotation or highlight
3. Annotations that you have created can be found under the highlights
4. Click on the **section title** to go directly to the page where an annotation or highlight is located.

The screenshot shows the ChemTours interface. On the left, there is a sidebar with an 'Edit note' section containing 15 notes and highlights. Below this is a 'List of ChemTours' section with a red '4' indicating four items. A context menu is open over the 'List of ChemTours' section, showing 'Edit' and 'Delete' options. The main content area displays a 'List of ChemTours' with a 'Previous: List of Applications' link. A teal banner at the top of the main content area reads 'ChemTours'. Below the banner, a list of topics is shown, including 'Dimensional Analysis', 'Significant Figures', 'Scientific Notation', 'Temperature Conversion', 'Cathode-Ray Tube', 'Millikan Oil-Drop Experiment', 'Rutherford Experiment', 'NaCl Reaction', 'Synthesis of Elements', and 'Avogadro's Number'.

Can I search my ebook for specific terms or page numbers?

Searching the Ebook

To search the text, select the magnifying glass from the left-hand side of the screen.

The screenshot shows the ChemTours interface with a search icon highlighted by a red arrow in the top-left corner. The page title is '3: Stoichiometry: Mass, Formulas, and Reactions'. The page number '82' is visible in the top-right corner. The main content area features a large '3' and the title 'Stoichiometry' followed by the subtitle 'Mass, Formulas, and Reactions'. Below the text is a photograph of two people in winter gear sitting on a snowy ground, looking at a large fire burning in a metal barrel.

Enter a term in the search field.

Chemistry

3: Stoichiometry: Mass, Formulas, and Reactions

Search

Atoms

Type in the field above to search the book

3

Stoichiometry

Mass, Formulas, and Reactions

See the full book search results displayed below.

Chemistry

3: Stoichiometry: Mass, Formulas, and Reactions

Search

Atoms

Cancel Search X

Brief Contents

1: Particles of Matter: Measurement and the Tools of Science

Questions and Problems

"...heterogeneous. (Section 1.2) LO2 All matter consists of **atoms**, and we use chemical formulas consisting of atomic..."

2: Atoms, Ions, and Molecules: Matter Starts Here

Questions and Problems

"...of atomic structure. (Sections 2.1 and 2.2) LO2 **Atoms** consist of a nucleus containing protons and neutrons..."

3: Stoichiometry: Mass, Formulas, and Reactions

Questions and Problems

"...in a balanced chemical equation, the number of **atoms** of each element is the same on the reactant side..."

4: Reactions in Solution: Aqueous

0 results in this section

3

Stoichiometry

Mass, Formulas, and Reactions

Clicking on the search results will take you to that specific page in the ebook. Additionally, the keyword you entered will appear highlighted in the text, and you will see a note at the top of the page indicating how many times that word is used within the section.

Chemistry

1: Particles of Matter: Measurement and the Tools of Science

Search

Atoms

Cancel Search X

Brief Contents

1: Particles of Matter: Measurement and the Tools of Science

Questions and Problems

"...heterogeneous. (Section 1.2) LO2 All matter consists of **atoms**, and we use chemical formulas consisting of atomic..."

2: Atoms, Ions, and Molecules: Matter Starts Here

Questions and Problems

"...of atomic structure. (Sections 2.1 and 2.2) LO2 **Atoms** consist of a nucleus containing protons and neutrons..."

3: Stoichiometry: Mass, Formulas, and Reactions

Questions and Problems

"...in a balanced chemical equation, the number of **atoms** of each element is the same on the reactant side..."

4: Reactions in Solution: Aqueous

0 of 6 results in this section

ANCIENT UNIVERSE The colors of the more than 10,000 galaxies in this image give us a glimpse into the universe as it existed about 13 billion years ago. This image was taken by NASA's Hubble Space Telescope.

PARTICULATE REVIEW

Atoms and Molecules: What's the Difference?

In Chapter 1 we explore how chemists classify different kinds of matter, from elements to compounds to mixtures. Hydrogen and helium were the first two elements formed after the universe began. Chemists use distinctively colored spheres to distinguish **atoms** of different elements in their drawings and models. For example, hydrogen is almost always depicted as white.

- How many of the following particles are shown in this image?
- Hydrogen **atoms**?
- Hydrogen molecules?
- Helium **atoms**?
- Are molecules composed of **atoms**, or are **atoms** composed of molecules?

SHOW ANSWER

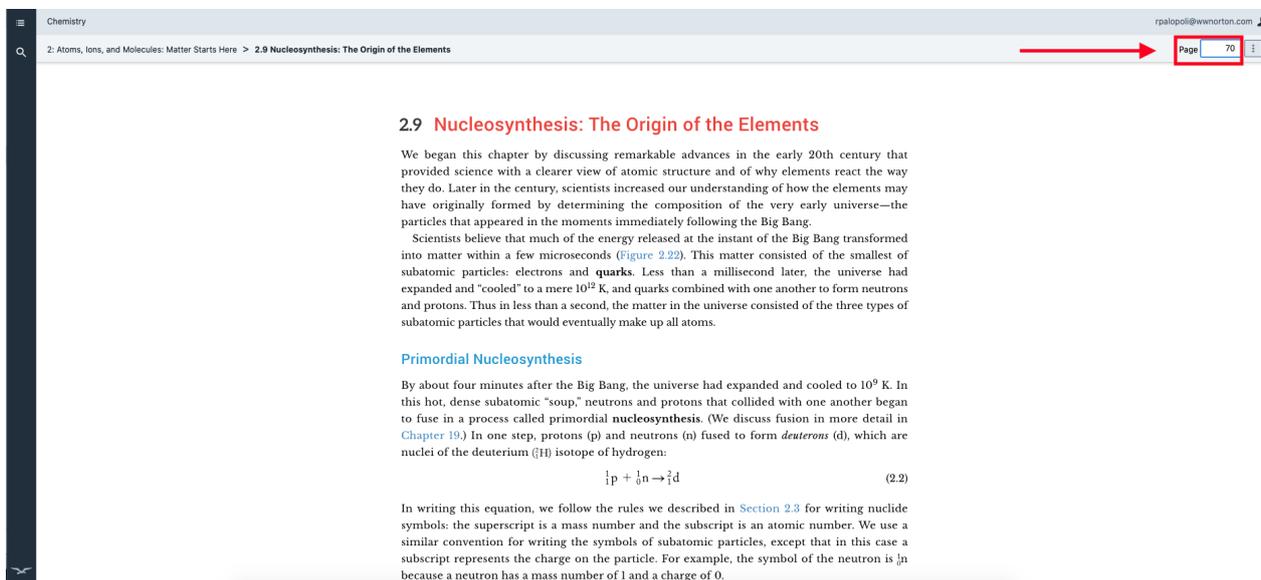
Learning Outcomes

Please Note: These search results are for the entire book. Anytime the keyword you entered is displayed in the text, it will show up here. If you would like to view the help notes on searching the

Table of Contents, please click here.

How do I search by page number?

You can search by a specific page number by inserting a number into the page field on the right-hand of the screen. This box will display the current page number you are viewing.



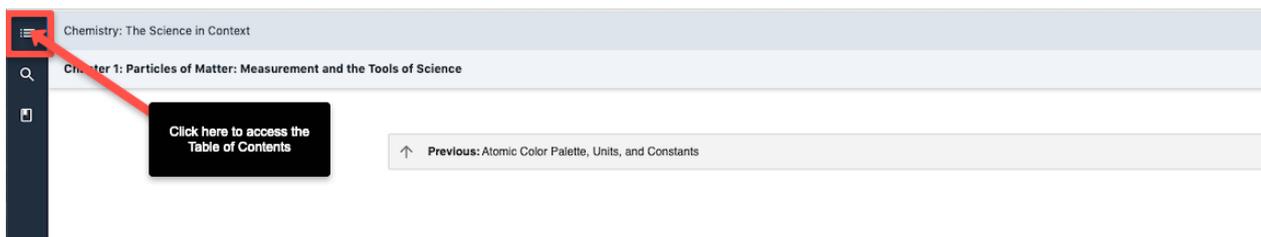
The screenshot shows the top navigation bar of an ebook reader. The breadcrumb trail is "2: Atoms, Ions, and Molecules: Matter Starts Here > 2.9 Nucleosynthesis: The Origin of the Elements". On the right side of the navigation bar, there is a "Page" field containing the number "70". A red arrow points from the text above to this field. The main content area displays the title "2.9 Nucleosynthesis: The Origin of the Elements" and the beginning of the text, including a section on "Primordial Nucleosynthesis" and a chemical equation:
$${}^1_1\text{p} + {}^1_0\text{n} \rightarrow {}^2_1\text{d} \quad (2.2)$$

After inserting a number in the page and select return on your keyboard, the ebook will take you to that page.

Can I read my ebook offline?

The newest version of the Norton Ebook Reader features the ability to read sections of your ebook offline.

Open any chapter of your ebook and select the **Table of Contents** icon.



The screenshot shows the top navigation bar of an ebook reader. The breadcrumb trail is "Chapter 1: Particles of Matter: Measurement and the Tools of Science". On the left side of the navigation bar, there is a "Table of Contents" icon (three horizontal lines). A red arrow points from the text above to this icon. A black callout box with white text says "Click here to access the Table of Contents". Below the navigation bar, there is a "Previous" button with an upward arrow and the text "Previous: Atomic Color Palette, Units, and Constants".

Select the 3 dot **Action Menu** from the ebook's Table of Contents view.

Chemistry: The Science in Context
Chapter 1: Particles of Matter: Measurement and the Tools of Science

Previous: Atomic Color Palette, Units, and Constants

1

Particles of Matter

Measurement and the Tools of Science

The screenshot shows the left sidebar of a digital textbook. The 'Table of Contents' is expanded to show Chapter 1: Particles of Matter: Measurement and the Tools of Science. A red box highlights a button labeled 'Select content for offline reading' next to the chapter title. A red arrow points from this button to the main content area on the right.

Click **Select content for offline reading**

Chemistry: The Science in Context
Chapter 3: Stoichiometry: Mass, Formulas, and Reactions

Previous: Summary

3

Stoichiometry

Mass, Formulas, and

The screenshot shows the left sidebar of a digital textbook. The 'Table of Contents' is expanded to show Chapter 3: Stoichiometry: Mass, Formulas, and Reactions. A red box highlights a button labeled 'Select content for offline reading' next to the chapter title. A red arrow points from this button to the main content area on the right.

Once offline reading is enabled, cached section **buttons** showing content available for offline reading will appear on the left as shown below.

Chemistry: The Science in Context

Chapter 3: Stoichiometry: Mass, Formulas, and Reactions

↑ Previous: Summary

3

Stoichiometry

Mass, Formulas, and Reactions



Table of Contents

Search Table of Contents

CHEMISTRY Chemistry: The Science in Context
Sixth Edition
by Natalie Foster

- Ⓢ Front Matter >
- Ⓢ Chapter 1: Particles of Matter: Measurement and the Tools of Science >
- Ⓢ Chapter 2: Atoms, Ions, and Molecules: Matter Starts Here >
- Ⓢ Chapter 3: Stoichiometry: Mass, Formulas, and Reactions >
- Ⓢ Chapter 4: Reactions in Solution: Aqueous Chemistry in Nature >
- Ⓢ Chapter 5: Properties of Gases: The Air We Breathe >
- Ⓢ Chapter 6: Thermochemistry: Energy Changes in Chemical Reactions >
- Ⓢ Chapter 7: A Quantum Model of Atoms: Waves, Particles, >

Select content for offline reading. **Cancel**

Select the content you would like to make available for offline reading by selecting the **button** to the left of the chapter. You can also use the arrows to the right of the chapter title to view more detailed options when selecting content. Once you have finished selecting content, a progress bar will appear at the bottom of the page.

Table of Contents Available Offline

Search Table of Contents

CHEMISTRY Chemistry: The Science in Context
Sixth Edition
by Natalie Foster

- Front Matter
- Chapter 1: Particles of Matter: Measurement and the Tools of Science
- Chapter 2: Atoms, Ions, and Molecules: Matter Starts Here
- Chapter 3: Stoichiometry: Mass, Formulas, and Reactions**
- Chapter 4: Reactions in Solution: Aqueous Chemistry in Nature
- Chapter 5: Properties of Gases: The Air We Breathe
- Chapter 6: Thermochemistry: Energy Changes in Chemical Reactions

Preparing for offline reading... (2/11 item)

3

Stoichiometry

Mass, Formulas, and Reactions



Please note: The more content you select, the longer it will take to make your selection available for offline reading.

When a section has been successfully cached, you will see the **Available Offline** tab as shown here and a check mark will appear next to the cached content.

The screenshot displays a digital chemistry textbook interface. On the left, a dark sidebar contains a 'Table of Contents' section with a search bar and a list of chapters. The chapter 'Chapter 3: Stoichiometry: Mass, Formulas, and Reactions' is highlighted with a red box, and a red arrow points to the 'Available Offline' tab above it. The main content area on the right shows the chapter title '3 Stoichiometry' and the subtitle 'Mass, Formulas, and Reactions'. Below the text is a photograph of a person in a laboratory setting, possibly a student, working with equipment. The interface includes navigation buttons like 'Previous: Summary' and 'Cancel'.

After the content you selected has been made available for offline reading, click on the **Available Offline** tab to view the sections cached for offline reading. The content can be accessed directly via the link as shown here:

Chemistry: The Science in Context

Chapter 3: Stoichiometry: Mass, Formulas, and Reactions

Table of Contents

Available Offline

Access your offline content directly with a link.

<https://nerd.wwnorton.com/nerd> Copy

Chapter 3: Stoichiometry: Mass, Formulas, and Reactions

Previous: Summary

3

Stoichiometry: Mass, Formulas, and Reactions

Select content for offline reading. Cancel

Click on the **Cancel** button to return to the Table of Contents.

Please Note: Media content (audio, video, animations, etc.) found in the ebook will only function with an active internet connection; those resources cannot be made available offline.

Highlights and Annotations

- In offline reading mode, you will only see notes and highlights for the content you've cached for offline reading.
- Notes and highlights cannot be created while in Offline Reading mode.

Parts of the ebook I previously made available offline are no longer available offline.

- Offline reading uses storage built into internet browsers. This means that when you make parts of your ebook available for offline reading, you aren't actually downloading or saving anything to your computer or mobile device.
- Sometimes, when your browsing history is cleared, the parts of the ebook you've made available offline will be cleared and will no longer be available for offline reading.
- To make these sections available for offline reading again, you'll need to get back online and repeat the same process you used to make those selections available offline the first time.

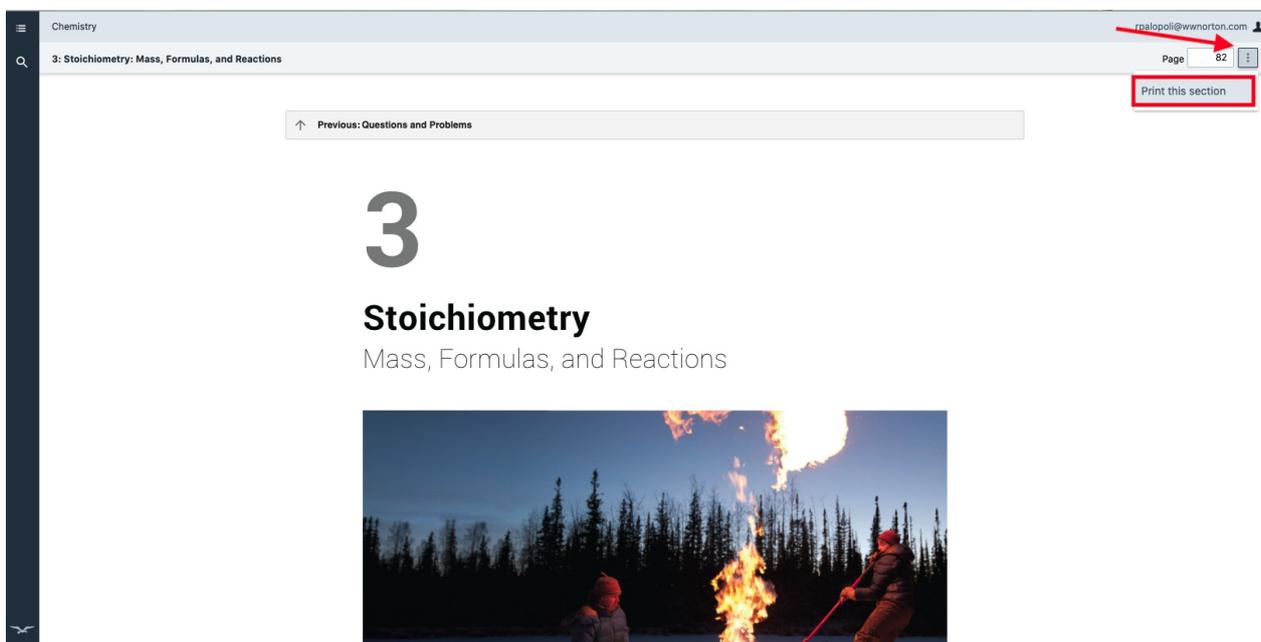
Does the ebook support audio narration?

Norton Ebooks and the Norton Ebook Reader are compatible both with screen readers and with browser extensions that enable text-to-speech functionality, such as the “Read Aloud” tool available on Google Chrome and Firefox. Apple and Google also offer text-to-speech solutions that you can use to read your ebook aloud on mobile devices. Please click on one of the links below for more information:

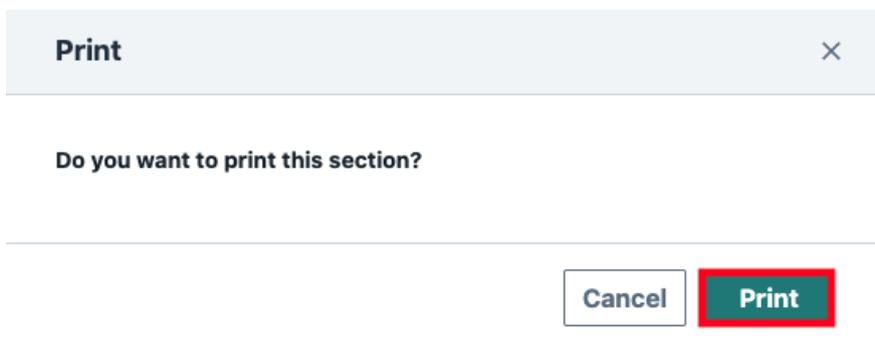
- Apple iOS: [VoiceOver](#)
- Chrome: [Read Aloud: A Text to Speech Voice Reader](#)
- Google Android: [Google Text-to-speech](#)
- Firefox: [Read Aloud: A Text to Speech Voice Reader](#)

How do I print a specific section of the ebook?

To print a specific section, select the three dots next to the page number at the top.

The screenshot shows the Norton Ebook Reader interface for the section '3: Stoichiometry: Mass, Formulas, and Reactions'. The page number '3' is prominently displayed. In the top right corner, there is a 'Page 62' indicator and a three-dot menu icon. A red box highlights the 'Print this section' button that appears next to the three-dot menu. Below the title, there is a large image of a bonfire in a snowy forest at night.

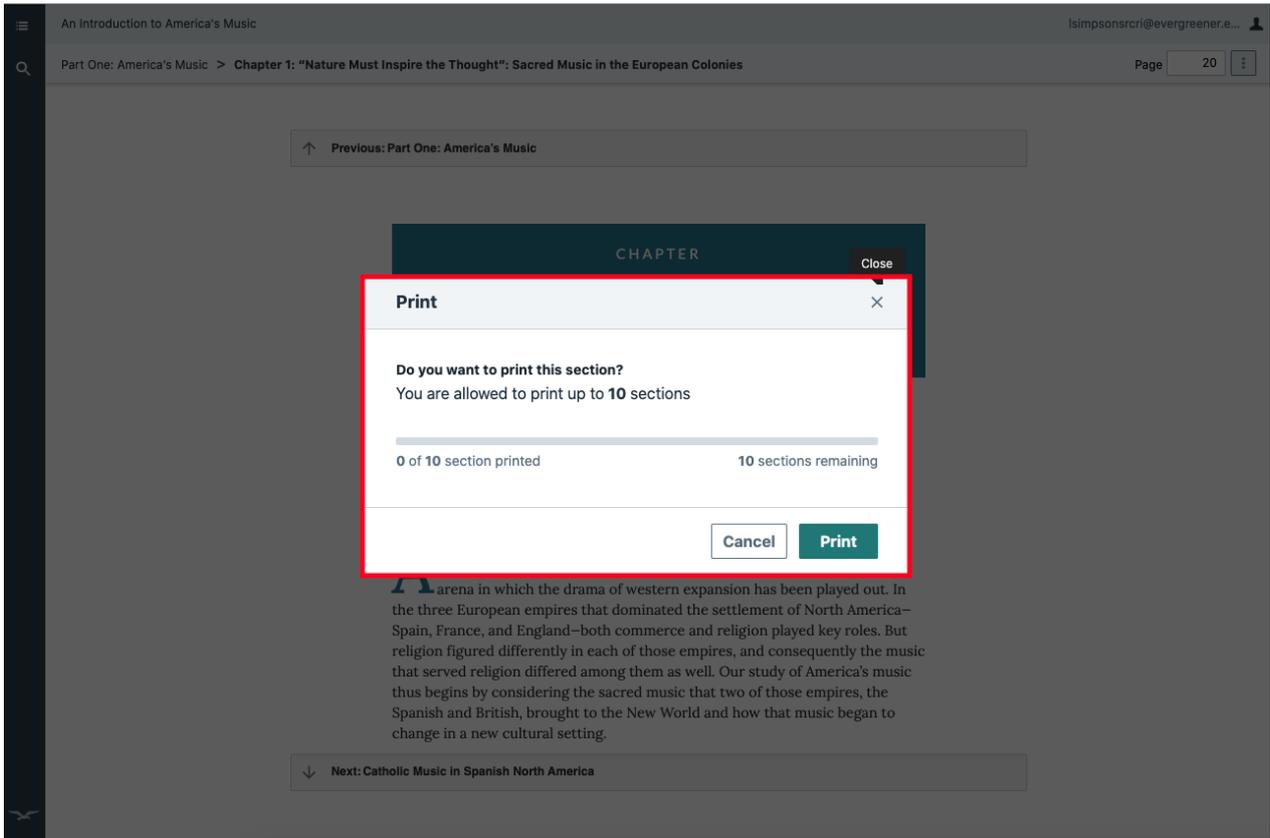
Select 'print' from the confirmation box.

The screenshot shows a print confirmation dialog box. At the top, it says 'Print' with a close button (X). Below that, it asks 'Do you want to print this section?'. At the bottom, there are two buttons: 'Cancel' and 'Print'. The 'Print' button is highlighted with a red border.

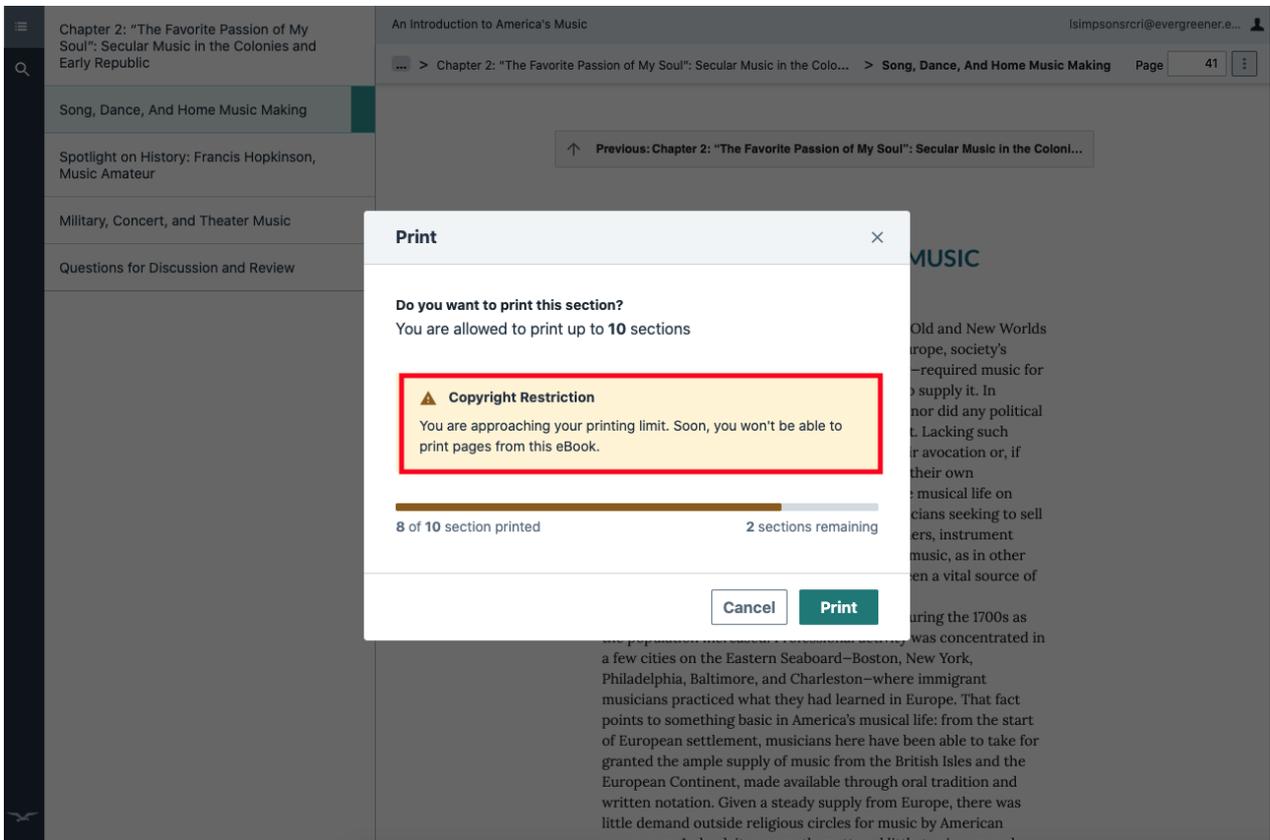
Can I print the entire ebook?

The number of sections you are allowed to print varies by title.

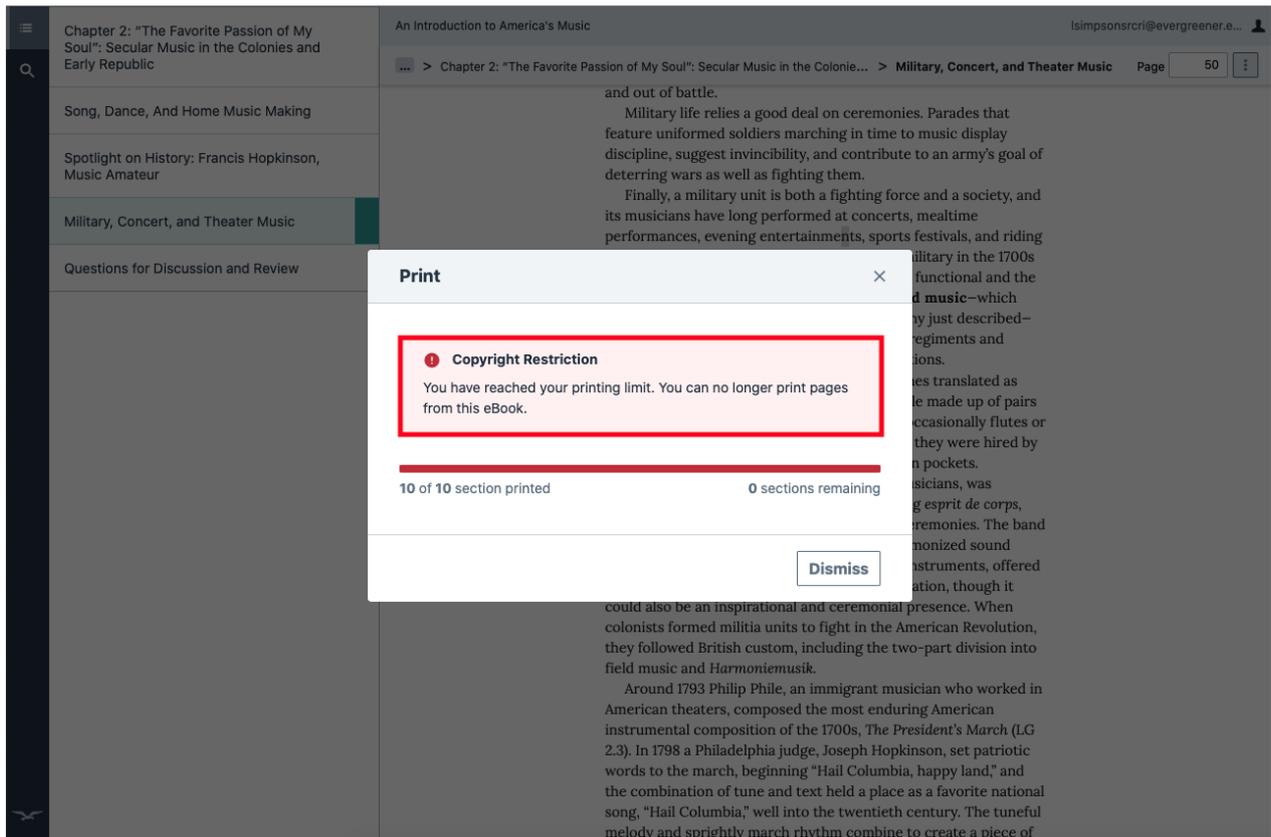
When you select print on a section, you will see a popup window with a progress bar which indicates how many sections you have already printed and how many you have left to print.



When you are close to the limit, you will see a copyright restriction message.



When you have reached the limit, you will see a message notifying you the limit has been reached.



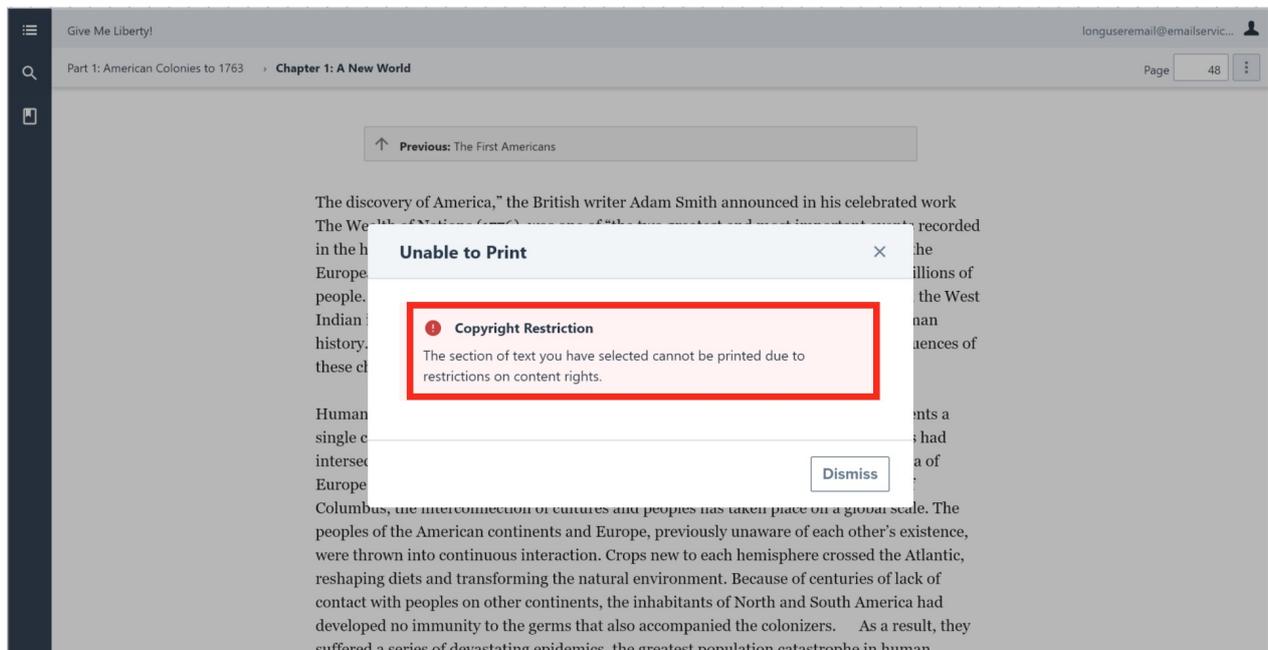
The screenshot shows a digital reading interface for an eBook titled "An Introduction to America's Music". The left sidebar contains a table of contents with sections like "Chapter 2: 'The Favorite Passion of My Soul': Secular Music in the Colonies and Early Republic", "Song, Dance, And Home Music Making", "Spotlight on History: Francis Hopkinson, Music Amateur", "Military, Concert, and Theater Music", and "Questions for Discussion and Review". The main content area displays text about military music, mentioning "Military life relies a good deal on ceremonies" and "Finally, a military unit is both a fighting force and a society". A "Print" dialog box is overlaid on the text, featuring a red border and a red icon. The message reads: "Copyright Restriction: You have reached your printing limit. You can no longer print pages from this eBook." Below the message is a progress bar showing "10 of 10 section printed" and "0 sections remaining". A "Dismiss" button is located at the bottom right of the dialog box.

Please Note: The print quota is based on copyright restrictions which are set by title. Print quotas cannot be reset on user accounts. Once you have printed the maximum amount allowed, you will not be able to print another section. Please keep this in mind when selecting sections to print.

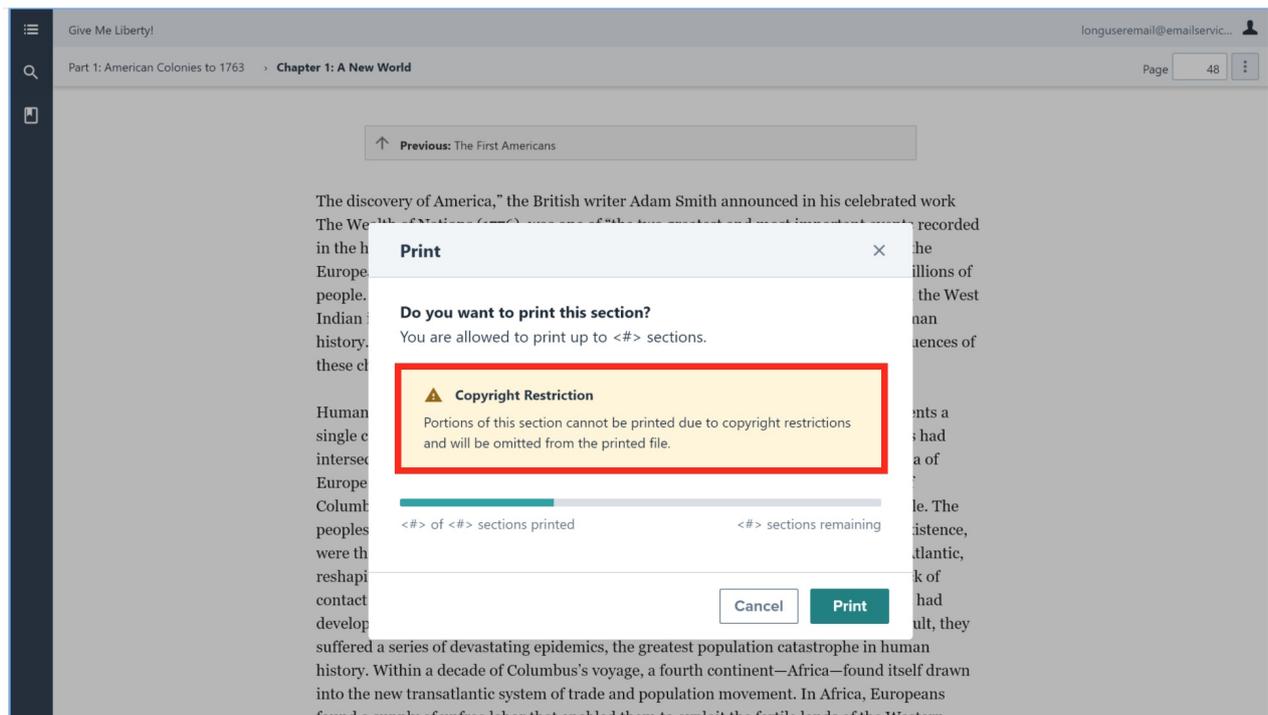
Why can't I print certain sections?

Some sections or parts of a section cannot be printed due to copyright restrictions. After selecting print this section, the popup you see will indicate if the section is available.

If the section you want to print is unavailable due to a copyright restriction, you will see the following message.



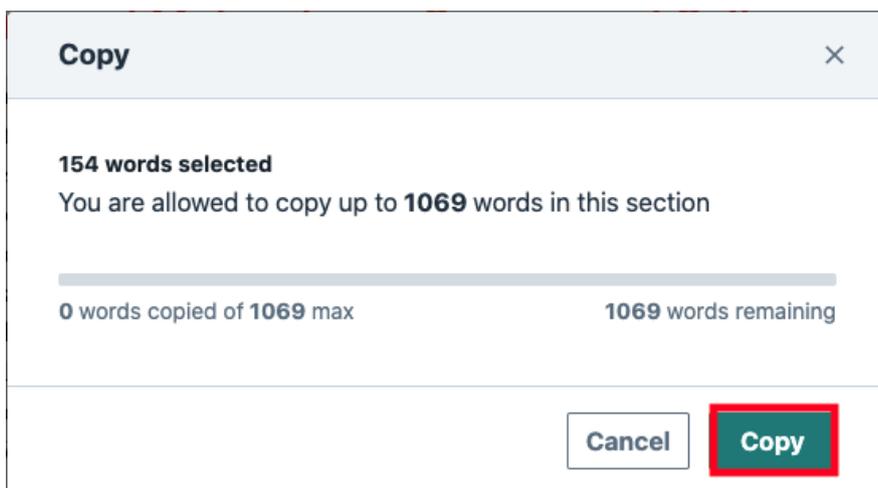
You may also see a message that indicates only part of the section is available for printing.



Can I copy and paste text from the ebook?

You can copy text by highlighting a portion of the text and using your keyboard shortcuts (command+c for mac, control+c for windows).

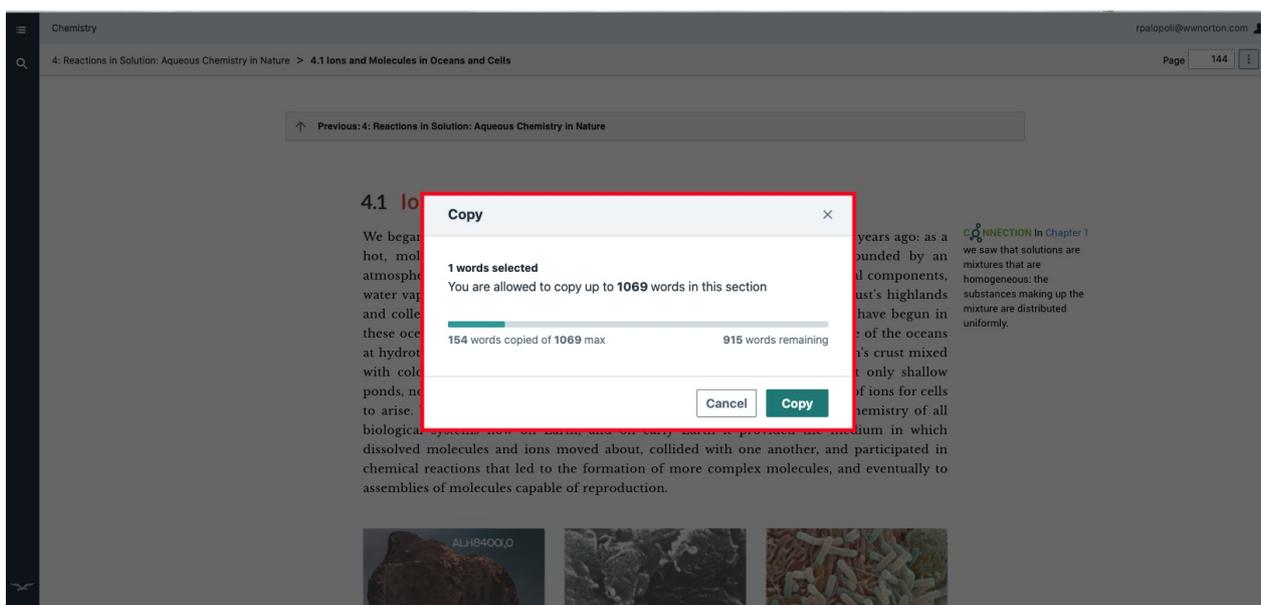
Once you have selected the area of the text you want to copy, you will see the following pop-up which indicates how many words you have copied, and how many you are allowed to copy from the ebook.



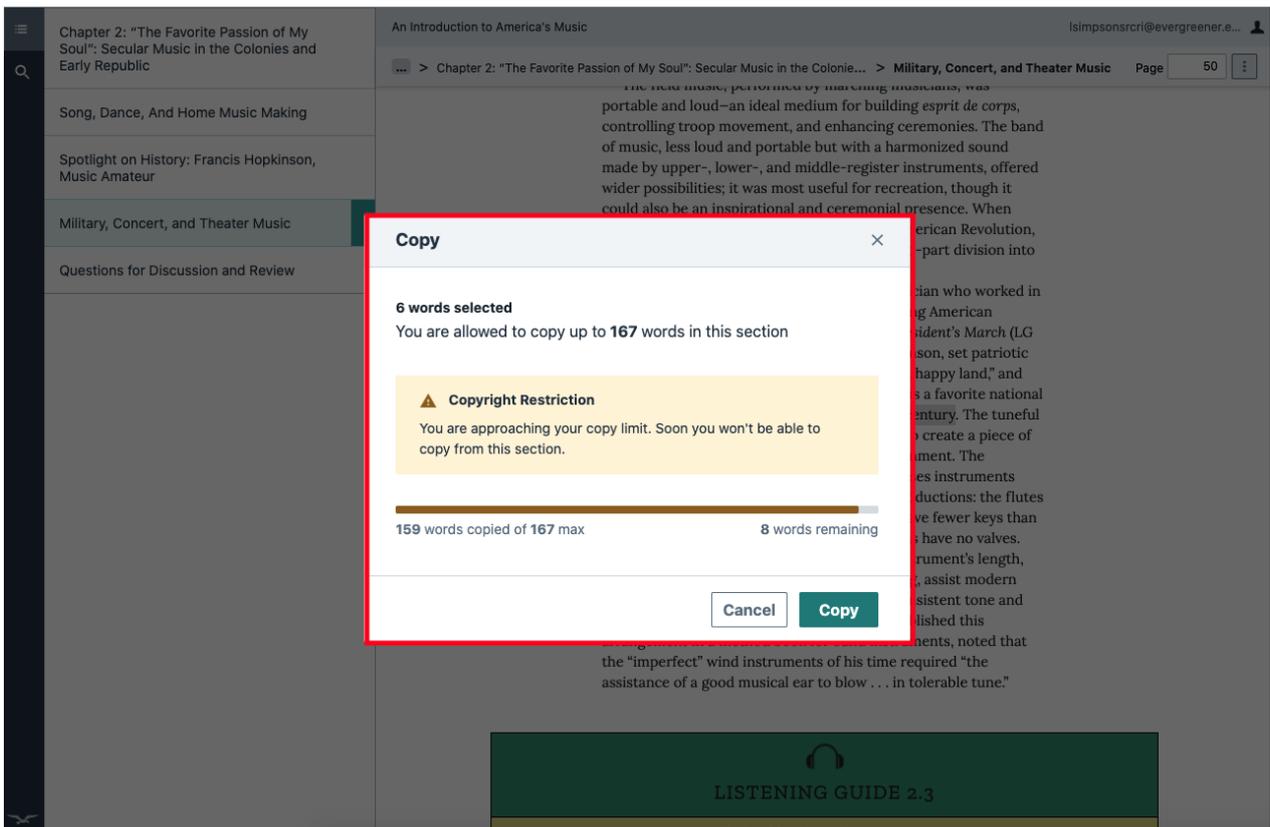
Can I copy the entire ebook?

The amount of content you can copy varies by title.

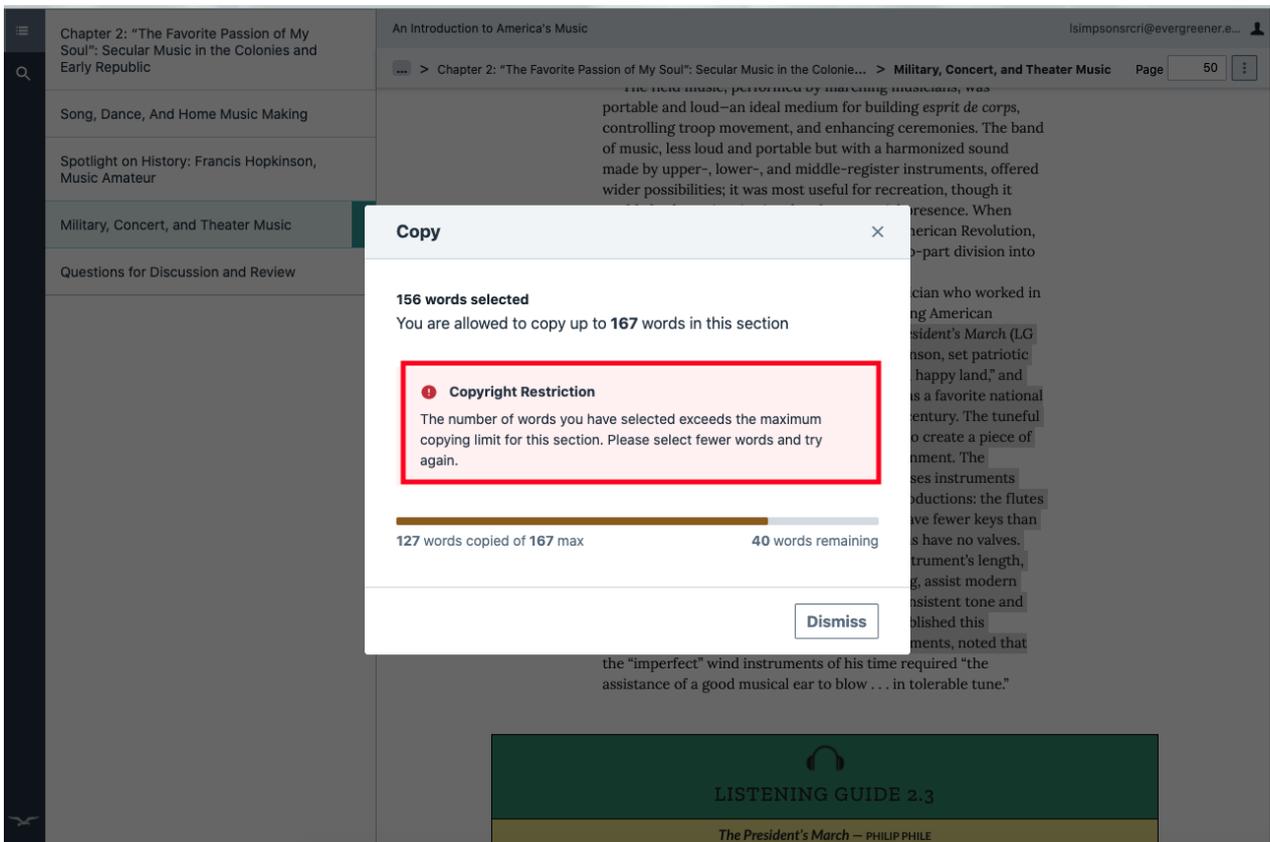
When you select a section of the text to copy, you will see a popup window with a progress bar indicating how many words you have already copied and how many you have left.



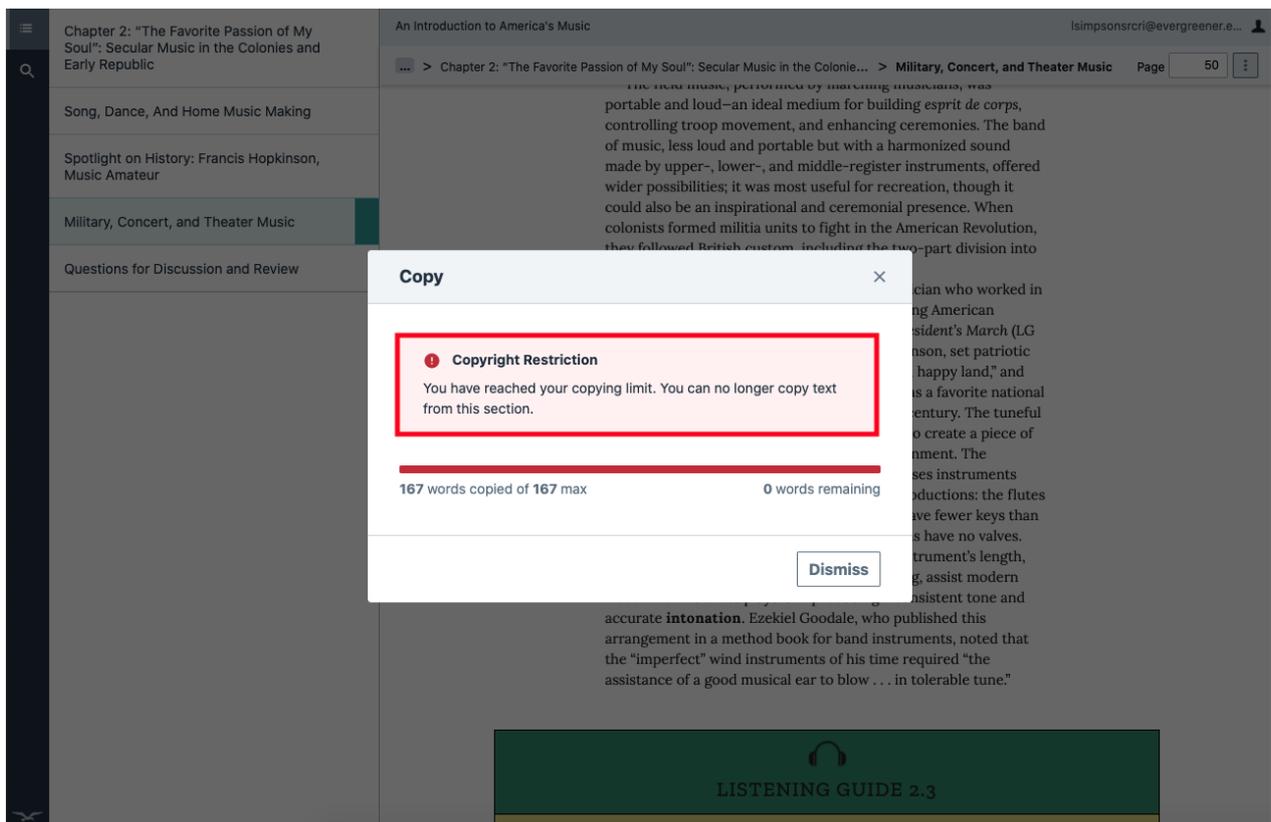
When you are approaching the limit, you will see a copyright restriction message.



If you have exceeded the copying limit, you will see the following message.



When you have reached the limit, you will see a message notifying you the limit has been reached.



Please Note: The copy quota is based on copyright restrictions which are set by title. Copy quotas cannot be reset on user accounts. Once you have copied the maximum amount allowed, you will not be able to copy another section. Please keep this in mind when selecting sections to copy.

Why can't I copy certain parts of the ebook?

Certain parts of the ebook cannot be copied due to copyright restrictions. After selecting an area to copy, the popup you see will indicate if the section is available.

If you are unable to copy the section, you will see the following message.

